

Chem 2430 HW #4 Answers

$$1. x = \sqrt{\frac{\hbar}{2m\omega}} (a + a^\dagger)$$

$$\begin{aligned} \langle 0 | x^4 | 0 \rangle &= \left(\frac{\hbar}{2m\omega} \right)^2 \langle 0 | (a + a^\dagger)^4 | 0 \rangle \\ &= \frac{\hbar^2}{4m^2\omega^2} \langle 0 | a a a^\dagger a^\dagger + a a^\dagger a a^\dagger | 0 \rangle \\ &= \frac{\hbar^2}{4m^2\omega^2} (2+1) = \frac{3}{4} \frac{\hbar^2 \omega^2}{k^2} \end{aligned}$$

$$\begin{aligned} \langle 0 | x^3 | 1 \rangle &= \left(\frac{\hbar}{2m\omega} \right)^{3/2} \langle 0 | (a + a^\dagger)^3 | 1 \rangle = \left(\frac{\hbar}{2m\omega} \right)^{3/2} \langle 0 | a a a^\dagger + a^\dagger a a | 1 \rangle \\ &= \left(\frac{\hbar}{2m\omega} \right)^{3/2} (2+1) \end{aligned}$$

$$\frac{|\langle 0 | x^3 | 1 \rangle|^2}{\hbar\omega} = \left(\frac{\hbar}{2m\omega} \right)^3 \frac{1}{\hbar\omega} 9 = \frac{9 \hbar^2}{8 m^3 \omega^4} = \frac{9}{8} \frac{\hbar^2 \omega^2}{k^3}$$

$$\begin{aligned} \langle 0 | x^3 | 3 \rangle &= \left(\frac{\hbar}{2m\omega} \right)^3 \langle 0 | (a + a^\dagger)^3 | 3 \rangle = \left(\frac{\hbar}{2m\omega} \right)^{3/2} \langle 0 | a a a | 3 \rangle \\ &= \left(\frac{\hbar}{2m\omega} \right)^{3/2} \sqrt{3} \sqrt{2} \end{aligned}$$

$$\frac{|\langle 0 | x^3 | 3 \rangle|^2}{3 \hbar \omega} = \frac{6}{3 \hbar \omega} \left(\frac{\hbar}{2m\omega} \right)^3 = \frac{1}{4} \frac{\hbar^2 \omega^2}{k^3}$$

The two x^3 terms combine to give $\frac{11}{8} \frac{\hbar^2 \omega^2}{k^3}$

The potential is described by $\frac{1}{2} k x^2 + \gamma x^3 + \delta x^4$

So the energy is given by

$$\frac{1}{2} \hbar \omega + \frac{3}{4} \frac{\hbar^2 \omega^2}{k^2} \gamma + \frac{11}{8} \frac{\hbar^2 \omega^2}{k^3} \delta$$

Note: if you wanted to calculate the energy of the $10 \rightarrow 11$ transition, you would need to also evaluate

$$\langle 1 | x^4 | 1 \rangle, \langle 1 | x^3 | 0 \rangle, \langle 1 | x^3 | 2 \rangle, \langle 1 | x^3 | 4 \rangle$$

2. $\hat{L}^2 \psi = 6\psi$ (we are assuming $\hbar=1$)

Since $\hat{L}^2 \psi = l(l+1)\psi$, this implies $l=2$

There are five orientations: $m_l = -2, -1, 0, 1, 2$.

3. $^1H^{3s}Cl$ $J=0 \rightarrow 1$ transition at 21.2 cm^{-1} .

$$E = J(J+1)B$$

$$E_0 \rightarrow E_1 = 2B$$

$$\text{So } B_e = 10.6 \text{ cm}^{-1}$$

$$\frac{1}{2\mu r^2} = B_e$$

In atomic units μ is 1.786×10^3 , $B_e = 4.829 \times 10^{-5}$

$$\text{So } r = 2.4077 \text{ au} \rightarrow 1.274 \text{ \AA}$$