

CHEM 1440

HW# 1

SOLUTIONS

1.3) From De Broglie Wavelength

$$\lambda = \frac{h}{mv} = \frac{h}{m\sqrt{\frac{3kT}{m}}} = \frac{h}{m^{1/2}\sqrt{3kT}} \Rightarrow T = \frac{h^2}{3km\lambda^2}$$

$$m_{He} = \frac{4.0 * 10^{-3} \text{ kg}}{6.02 * 10^{23}} = 6.64 * 10^{-27} \text{ kg He atom}$$

$$T_{He} = \frac{h^2}{3km\lambda^2} = \frac{(6.626 * 10^{-34} \text{ Js})^2}{3 * (0.2 * 10^{-9} \text{ m})^2 * 1.38 * 10^{-23} \text{ J} * \text{K}^{-1} * 6.64 * 10^{-27} \text{ kg}} \approx 40 \text{ K}$$

$$m_{Ar} = 6.63 * 10^{-26} \text{ kg Ar atom}$$

$$T_{Ar} \approx 4 \text{ K}$$

1.6)

a) $E = 0.10 \text{ J}$ and $\Delta t = 10 \text{ ns} = 10 * 10^{-9} \text{ s} = 10^{-8} \text{ s}$

$$P = E / \Delta t = 10^7 \text{ Js}^{-1} = 10^7 \text{ W}$$

b)

$$E_{\text{photon}} = \frac{hc}{\lambda} = \frac{(6.626 * 10^{-34} \text{ Js})(3 * 10^8 \text{ m/s})}{(1000 * 10^{-9} \text{ m})} \approx 1.99 * 10^{-19} \text{ J}$$

$$n_{\text{photon}} = \frac{E}{E_{\text{photon}}} = \frac{0.1 \text{ J}}{1.99 * 10^{-19} \text{ J}} \approx 5.0 * 10^{17} \text{ photons}$$

1.11)

Work Function (I_{WF}) & Planck's Constant (h)

$$E_{\text{Kinetic}} = E_{\text{photon}} - I_{WF}$$

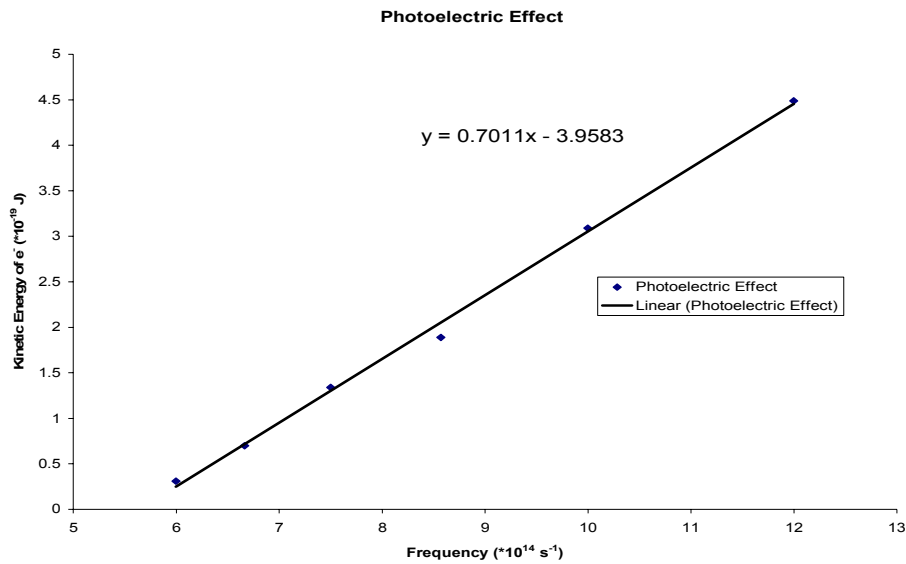
**** Kinetic Energy given in the question should be 10^{-19} J not 10^{19} J .**

Kinetic Energy (10^{-19} J)	4.49	3.09	1.89	1.34	0.700	0.311
Wavelength (nm)	250	300	350	400	450	500
Frequency (10^{14} s^{-1}) ($\nu = hc/\lambda$)	12	10	8.57	7.5	6.67	6.0

From the best line (E_{kinetic} vs the frequency of photons)

$$E_{kinetic} = 0.7011\nu - 3.96;$$

$$\rightarrow I_{WF} = 3.96 * 10^{-19} \text{ J} \approx 2.5 \text{ eV} \ \& \ h = 0.7011 * 10^{-19} \text{ J}/10^{14} \text{ s}^{-1} \approx 7.0 * 10^{-34} \text{ Js}$$



1.21)

$$I_{WF} = 5.65 \text{ eV} = 9.05 * 10^{-19} \text{ J}$$

$$E_{Kinetic} = E_{photon} - I_{WF} ;$$

$$\text{For minimum frequency} \rightarrow E_{Kinetic} = 0 ;$$

$$\rightarrow E_{photon} = I_{WF} = 9.05 * 10^{-19} \text{ J}$$

$$h\nu_{\min} = 9.05 * 10^{-19} \text{ J} \Rightarrow \nu_{\min} = \frac{9.05 * 10^{-19} \text{ J}}{6.626 * 10^{-34} \text{ Js}} = 1.37 * 10^{15} \text{ s}^{-1}$$

$$\text{If } \lambda = 150 \text{ nm then } E_{photon} = \frac{hc}{\lambda} = \frac{(6.626 * 10^{-34} \text{ Js})(3.0 * 10^8 \text{ ms}^{-1})}{150 * 10^{-9} \text{ m}} = 13.2 * 10^{-19} \text{ J}$$

$$E_{Kinetic} = E_{photon} - I_{WF} = 4.20 * 10^{-19} \text{ J}$$

$$\frac{1}{2} m_e v_e^2 = 4.20 * 10^{-19} \text{ J}$$

$$v_e^2 = \frac{2 * 4.20 * 10^{-19} \text{ J}}{m_e} = \frac{8.40 * 10^{-19} \text{ J}}{9.1 * 10^{-31} \text{ kg}} = 9.23 * 10^{11} \text{ m}^2 \text{ s}^{-2}$$

$$v_e = 9.6 * 10^5 \text{ ms}^{-1}$$